

(19)



Europäisches Patentamt

European Patent Office

Office européen des brevets



(11)

EP 0 861 817 A1

(12)

EUROPEAN PATENT APPLICATION

published in accordance with Art. 158(3) EPC

(43) Date of publication:

02.09.1998 Bulletin 1998/36

(51) Int. Cl.⁶: **C06D 5/00**, C06B 25/00,
C06B 31/12, B60R 21/26

(21) Application number: 96938456.9

(86) International application number:
PCT/JP96/03327

(22) Date of filing: 13.11.1996

(87) International publication number:
WO 97/18178 (22.05.1997 Gazette 1997/22)

(84) Designated Contracting States:
DE ES FR GB IT

(72) Inventor: **YAMATO, Yo**
Himeji-shi Hyogo 671-12 (JP)

(30) Priority: 14.11.1995 JP 295148/95

(74) Representative:
Grünecker, Kinkeldey,
Stockmair & Schwanhäusser
Anwaltssozietät
Maximilianstrasse 58
80538 München (DE)

(71) Applicant:
DAICEL CHEMICAL INDUSTRIES, LTD.
Sakai-shi, Osaka 590 (JP)

(54) GAS GENERATING COMPOSITION

(57) A gas generating composition having a low combustion temperature and a low heat of combustion at the time of burning and therefore enabling downsizing of gas generators themselves, which comprises as the essential components nitroguanidine and an oxidizing agent comprising (a) nitrates or nitrites of alkali metals or alkaline earth metals, (b) oxides or multiple oxides of metals selected from among copper, cobalt, iron, manganese, nickel, zinc, molybdenum and bismuth, or a mixture of the components (a) and (b).

EP 0 861 817 A1

Description

[Technical Field to which the Invention Belongs]

5 The present invention relates to a gas generating composition which serves a working gas in the air bag system to be mounted in automobiles, aircraft or the like for the protection of the human body.

[Prior Art]

10 Sodium azide is known to the public as the gas generating agent to be used in the air bag system for automobiles, aircraft or the like. Further, gas generating compositions comprising sodium azide are not particularly problematic in the burning characteristics and therefore are widely put to practical use. However, sodium azide has essentially unfavorable disadvantages. For example, many patents of this field have pointed out various problems such as dangerousness of decomposition explosion, formation of explosive compounds by the reaction with heavy metals, environmental pollution caused by mass disposal of the waste and so on.

15 Meanwhile, investigations are in progress on the substitutes for sodium azide to solve these problems. For example, gas generating compositions comprising transition metal complexes of tetrazole or triazole are disclosed in JP-B 6-57629; ones comprising triaminoguanidine nitrate in JP-A 5-254977; ones comprising carbonylhydrazide in JP-A 6-239683; and ones comprising cellulose acetate and nitrogenous nonmetallic compounds such as nitroguanidine in JP-
20 A 7-61885.

Further, USP 5125684 discloses the use of nitroguanidine as the energetic component to be made coexistent with 15 to 30% cellulose-based binder; and WO-A 96/25375 (published on August 22, 1996) discloses combinations of nitroguanidine with silicon oxide or aluminum oxide.

25 When the burning of a nitrogenous organic compound is conducted by the use of an oxidizing agent in a stoichiometric amount, i.e., such an amount as to generate enough oxygen to complete the burning of carbon, hydrogen and other elements constituting the compound, the heat of combustion and the combustion temperature are generally higher than those found in the burning of azide compounds disadvantageously. In addition to the performance of a gas generating agent, it is essential to the inflator system for air bags that the system itself has such a size as not to hinder the ordinary driving of an automobile. When a gas generating agent having a high combustion temperature and a high
30 heat of combustion is used, however, the use of additional parts for heat removal is unavoidable in designing a gas generator, which makes it impossible to downsize a gas generator. In short, it is most preferable that a gas generating agent have a low combustion temperature and a low heat of combustion at the time of burning. Accordingly, it cannot be said that the known gas generating compositions described above are satisfactory in applicability to air bag systems.

35 [Disclosure of the Invention]

The inventors of the present invention have intensively studied to solve the above problems, and the present invention has been accomplished as a result of the studies.

40 The present invention relates to a gas generating composition which comprises as the essential components nitroguanidine and an oxidizing agent comprising the following component (a) or (b) or a mixture of the components (a) and (b):

(a) nitrates or nitrites of alkali metals or alkaline earth metals, and

45 (b) an oxide or a multiple oxide of a metal(s) selected from copper, cobalt, iron, manganese, nickel, zinc, molybdenum and bismuth.

The composition is preferably a combination of nitroguanidine with the component (b), or with the components (a) and (b).

50 The composition preferably comprises 20 to 80% by weight of nitroguanidine and 80 to 20% by weight of the oxidizing agent, and still preferably comprises 25 to 40% by weight of nitroguanidine and 75 to 60% by weight of the oxidizing agent.

The composition may further contain at most 5% by weight of a binder or binders based on the total weight of the composition.

55 The oxidizing agent is preferably one wherein the component (a) is an alkaline earth metal nitrate, still preferably strontium nitrate, one wherein the component (a) is an alkali metal nitrate, still preferably potassium nitrate, one wherein the component (b) is an oxide, still preferably copper oxide, or one comprising strontium nitrate and copper oxide.

The modes for carrying out the invention will now be described in detail.

The content of nitroguanidine in the gas generating composition of the present invention is preferably 20 to 80% by

weight, still preferably 20 to 60% by weight, though it depends on the kind of the oxidizing agent and the oxygen balance. The content is most desirably 25 to 40% by weight because, in such a case, a gas generating composition comprising a copper oxide as the oxidizing agent exhibits a combustion temperature of 2200 K or below and a gas yield of 1.2 mol/100 g or above.

In the present invention, the above component (a) or (b) or a mixture of both is used as the oxidizing agent. Specific examples of (a) nitrites or nitrates of alkali metals or alkaline earth metals include alkali metal or alkaline earth metal salts of nitric acid, such as sodium nitrate, potassium nitrate, magnesium nitrate, strontium nitrate; and alkali metal or alkaline earth metal salts of nitrous acid, such as sodium nitrite, potassium nitrite, magnesium nitrite and strontium nitrite. On the other hand, specific examples of (b) oxides or multiple oxides of metals selected from copper, cobalt, iron, manganese, nickel, zinc, molybdenum and bismuth include CuO, Cu₂O, Co₂O₃, CoO, Co₃O₄, Fe₂O₃, FeO, Fe₃O₄, MnO₂, Mn₂O₃, Mn₃O₄, NiO, ZnO, MoO₃, CoMoO₄, Bi₂MoO₆ and Bi₂O₃. Among these oxidizing agents, copper oxides are the most desirable.

The oxidizing agent to be used in the present invention may be any one selected from among the above compounds or any combination of two or more members selected from among them. The content of the oxidizing agent in the gas generating composition is preferably 80 to 20% by weight, still preferably 80 to 40% by weight. The content is most desirably 75 to 60% by weight because, in such a case, a gas generating composition comprising a copper oxide as the oxidizing agent exhibits a combustion temperature of 2200 K or below and a gas yield of 1.2 mol/100 g or above.

The gas generating composition of the present invention may further contain a binder or binders. The binder to be used in the present invention includes inorganic ones such as silica, alumina and molybdenum disulfide; and organic ones such as microcrystalline cellulose, POVAL and high-molecular-weight oligomers. It is preferable that the binder content of the composition be 5% by weight or below.

The gas generating composition of the present invention comprising nitroguanidine and the above oxidizing agent generally exhibits a lowered combustion temperature and a lowered total heat of combustion as compared with those of the gas generating compositions of the prior art. In particular, a composition comprising nitroguanidine and CuO is extremely excellent in these respects, thus being the most important combination. Further, this composition can change in the burning velocity and the pressure exponent by controlling the oxygen balance. The term "pressure exponent" as used in this description refers to the exponent "n" in the formula: burning velocity r (mm/sec) = $a \times p^n$ (wherein a is a constant dependent on the constituents of a gas generating composition and initial temperature; and p is pressure (kgf/cm²) of measurement). The above composition has the property that n approaches 0 as the oxygen balance shifts to the plus side, which is particularly preferable from the standpoint of the stability of combustion.

The gas generating composition of the present invention is preferably prepared by mixing the components in a powdery state, and the mixing can be conducted also according to a wet process in the presence of water or the like at need. Prior to the use, the composition may, if necessary, be molded into a suitable shape such as granule, pellet or disk. Further, a composition having a low burning velocity may be molded by extrusion prior to the use.

The gas generating composition of the present invention is particularly useful for the air bag system to be mounted in automobiles, aircraft or the like for the protection of the human body.

The nitroguanidine contained in the gas generating composition of the present invention exhibits high long-term stability required of the air bag system and is excellent in burning characteristics.

[Example]

The present invention will now be described specifically by referring to the following Examples and Comparative Examples, though the present invention is not limited by them. In the following Examples and Comparative Examples, the compositions were pelletized in an ordinary manner before being examined.

Examples 1 to 5 and Comparative Examples 1 to 7

Table 1 shows the theoretical combustion temperatures of nitroguanidine-containing gas generating compositions according to the present invention. And as Comparative Examples are shown the theoretical combustion temperatures of gas generating compositions (Comparative Examples 1, 2) comprising transition metal complexes of 5-aminotetrazole (5-AT) (disclosed in JP-B 6-57629), gas generating composition (Comparative Example 3) comprising triaminoguanidine nitrate (disclosed in JP-A 5-254977), gas generating composition (Comparative Example 4) containing carbonylhydrazide (disclosed in JP-A 6-239683), and gas generating compositions (Comparative Examples 5, 6 and 7) comprising cellulose acetate and nitrogenous nonmetallic compounds (disclosed in JP-A 7-61885).

Examples 6 to 14

A gas generating composition comprising nitroguanidine and CuO could change in combustion temperature, burn-

ing velocity, density of pellets of the gas generating composition and gas yield, by changing the mixing ratio of the components. The data are given in Table 2. Each burning velocity was determined under a pressure of 70 kgf/cm².

As will be understood from the results, the gas generating compositions of the present invention are superior to those of the prior art particularly in combustion temperature, which paves the way to downsizing of gas generators and application to air bag systems.

Table 1

		Compn.	(wt %)	Combustion temp. (K)
Ex.	1	nitroguanidine/KNO ₃	(56.3/43.7)	2200
	2	nitroguanidine/Sr(NO ₃) ₂ /CuO	(40.3/19.2/40.5)	2091
	3	nitroguanidine/CuO	(39.5/60.5)	2043
	4	nitroguanidine/KNO ₃ /Al ₂ O ₃	(55.2/42.8/2.0)	2172
	5	nitroguanidine/CuO/cellulose	(32.8/64.7/2.5)	1928
Comp. Ex.	1	Zn(5-AT) ₂ /Sr(NO ₃) ₂	(44.0/56.0)	2411
	2	[Cu(S-AT) _{2.1} /2H ₂ O]/Sr(NO ₃) ₂	(42/58)	2390
	3	triaminoguanidine nitrate/KClO ₄	(57.9/42.1)	2911
	4	carbohydrazide/KClO ₄ /CaO	(39/61/10)	2825
	5	cellulose acetate/triacetin/KClO ₄ /nitroguanidine	(8/2/55/35)	2834
	6	cellulose acetate/triacetin/KClO ₄ /triaminoguanidine nitrate	(8/4/57/31)	2893
	7	cellulose acetate/triacetin/KClO ₄ /5-aminotetrazole	(10/5/65/20)	2928

Table 2

		Compn.	(wt %)	Combustion temp. (K)	Burning velocity (mm/s)	Density (g/cm ³)	Gas yield (mol/100 g of gas generating compn.)
Ex.	6	nitroguanidine/CuO	(39.5/60.5)	2043	4.3	2.54	1.90
	7	nitroguanidine/CuO	(38/62)	1992	4.9	2.56	1.83
	8	nitroguanidine/CuO	(36/64)	1922	5.8	2.58	1.73
	9	nitroguanidine/CuO	(34/66)	1850	6.9	2.63	1.63
	10	nitroguanidine/CuO	(32/68)	1774	8.8	2.71	1.54
	11	nitroguanidine/CuO	(30/70)	1695	9.1	2.76	1.44
	12	nitroguanidine/CuO	(28/72)	1604	10.6	2.83	1.35
	13	nitroguanidine/CuO	(26/74)	1526	11.0	2.94	1.25
	14	nitroguanidine/CuO	(24.6/75.4)	1517	9.5	2.94	1.18

Claims

1. A gas generating composition characterized by comprising as the essential components nitroguanidine and an oxidizing agent comprising the following component (a) or (b) or a mixture of the components (a) and (b):

- (a) nitrates or nitrites of alkali metals or alkaline earth metals, and
- (b) oxides or multiple oxides of metals selected from copper, cobalt, iron, manganese, nickel, zinc, molybdenum and bismuth.

EP 0 861 817 A1

2. The composition according to claim 1, which comprises nitroguanidine and the component (b).
3. The composition according to claim 1, which comprises nitroguanidine and the components (a) and (b).
- 5 4. The composition according to claim 1, which comprises 20 to 80% by weight of nitroguanidine and 80 to 20% by weight of the oxidizing agent.
- 10 5. The composition according to claim 1, which comprises 25 to 40% by weight of nitroguanidine and 75 to 60% by weight of the oxidizing agent.
6. The composition according to claim 1, which further contains a binder in an amount of 5% by weight or below based on the total amount of the composition.
7. The composition according to claim 1, wherein the component (a) is an alkaline earth metal nitrate.
- 15 8. The composition according to claim 1, wherein the component (a) is strontium nitrate.
9. The composition according to claim 1, wherein the component (a) is an alkali metal nitrate.
- 20 10. The composition according to claim 1, wherein the component (a) is potassium nitrate.
11. The composition according to claim 1, wherein the component (b) is an oxide.
12. The composition according to claim 1, wherein the component (b) is a copper oxide.
- 25 13. The composition according to claim 1, wherein the oxidizing agent comprises strontium nitrate and a copper oxide.

INTERNATIONAL SEARCH REPORT

International application No.

PCT/JP96/03327

A. CLASSIFICATION OF SUBJECT MATTER Int. Cl ⁶ C06D5/00, C06B25/00, 31/12, B60R21/26 According to International Patent Classification (IPC) or to both national classification and IPC		
B. FIELDS SEARCHED Minimum documentation searched (classification system followed by classification symbols) Int. Cl ⁶ C06D5/00, C06B25/00, 31/12, B60R21/26 Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched Jitsuyo Shinan Koho 1922 - 1996 Jitsuyo Shinan Toroku Kokai Jitsuyo Shinan Koho 1971 - 1997 Koho 1996 - 1997 Toroku Jitsuyo Shinan Koho 1994 - 1997 Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)		
C. DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
P, X	JP, 8-231291, A (Daicel Chemical Industries, Ltd.), September 10, 1996 (10. 09. 96), Claim; paragraphs (0009) to (0014), (0019); examples 20 to 25; comparative examples 11 to 13 & DE, 19548917, A1	1 - 13
X	JP, 6-227884, A (Nippon Koki Co., Ltd. and two others), August 16, 1994 (16. 08. 94), Claim; paragraphs (0018) to (0021), (0023); examples 1 to 2 & CN, 1083463, A	1 - 13
A	JP, 7-61885, A (Nippon Oil and Fats Co., Ltd.), March 7, 1995 (07. 03. 95), Claim & DE, 4412871, A1 & FR, 2703990, A1 & US, 5482579, A	1 - 13
<input type="checkbox"/> Further documents are listed in the continuation of Box C. <input type="checkbox"/> See patent family annex.		
* Special categories of cited documents: "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art "&" document member of the same patent family		
Date of the actual completion of the international search February 10, 1997 (10. 02. 97)		Date of mailing of the international search report February 18, 1997 (18. 02. 97)
Name and mailing address of the ISA/ Japanese Patent Office Facsimile No.		Authorized officer Telephone No.

Form PCT/ISA/210 (second sheet) (July 1992)